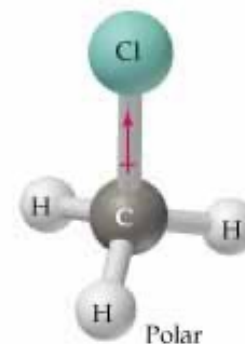


DIPOLE MOMENTS

What are polar and nonpolar covalent bonds, and how are they formed?

Dipole Moments

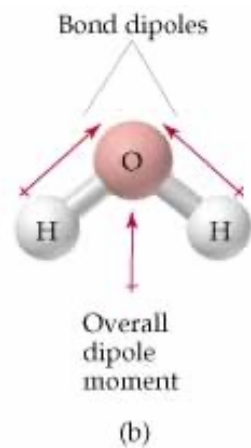
- Represents the atom's polarity magnitude and direction (+ \rightarrow -)
- Examples: H – Cl, NH₃, CH₃Cl, (& –COOH)



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Dipole Moments

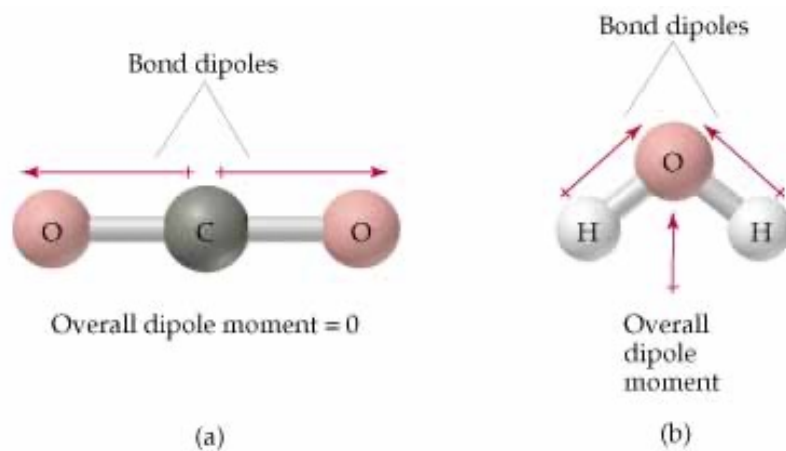
- Represents the atom's polarity magnitude and direction (+ → -)
- Examples: H – Cl, NH₃, CH₃Cl, (& –COOH)
- Water:



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Dipole Moments

- Represents the atom's polarity magnitude and direction (+ \rightarrow -)
- Bond Polarity and Shape Matters:



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