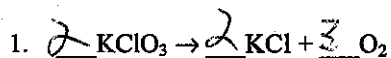


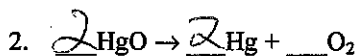
CHEMICAL REACTIONS REVIEW SHEET

BALANCE THE FOLLOWING EQUATIONS:

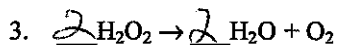


Name: Key

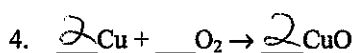
Type: decomposition



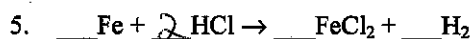
Type: decomposition



Type: decomposition



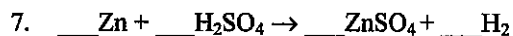
Type: Synthesis



Type: single displacement



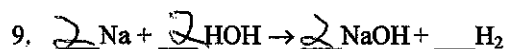
Type: Synthesis



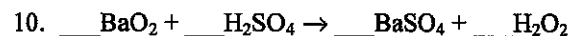
Type: single displacement



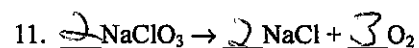
Type: Synthesis



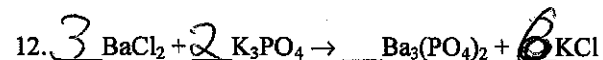
Type: single displacement



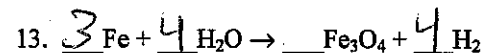
Type: double displacement



Type: decomposition



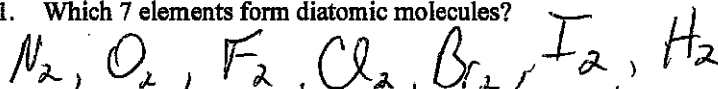
Type: double displacement



Type: single displacement

QUESTIONS:

1. Which 7 elements form diatomic molecules?

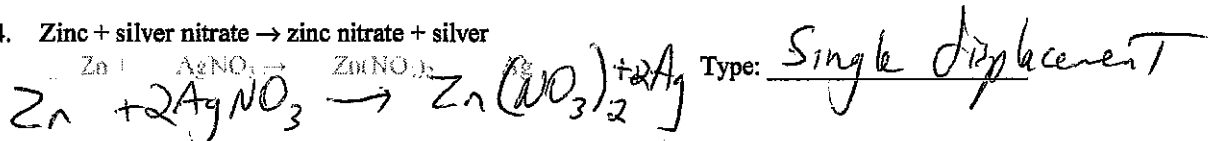


2. The products of neutralization reactions are a acid and base (OH compound)

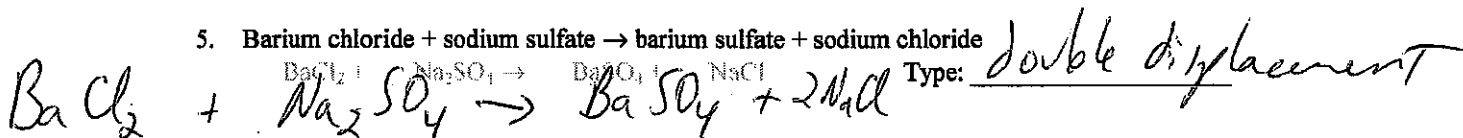
3. Combustion reactions take place between certain carbon substances and oxygen gas.

COMPLETE BALANCED FORMULA EQUATION ON THE SPACE BENEATH EACH OF THE FOLLOWING WORD EQUATIONS.

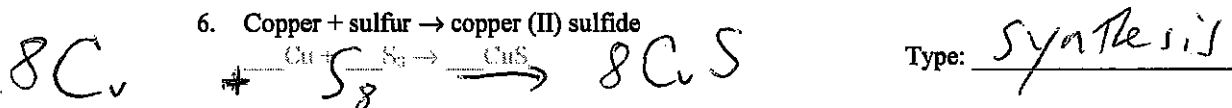
4. Zinc + silver nitrate → zinc nitrate + silver



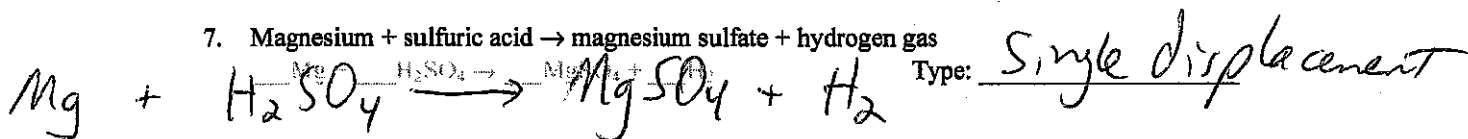
5. Barium chloride + sodium sulfate → barium sulfate + sodium chloride



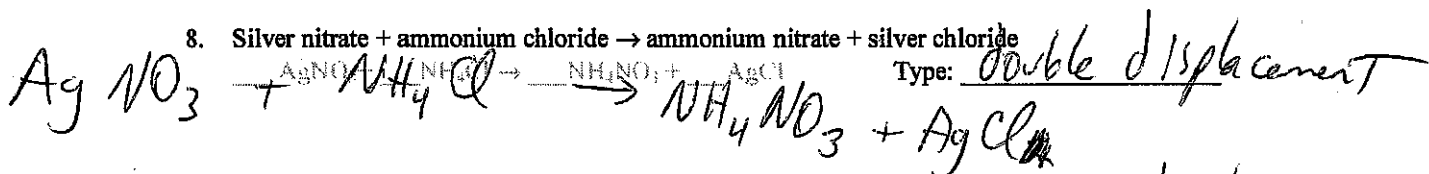
6. Copper + sulfur → copper (II) sulfide



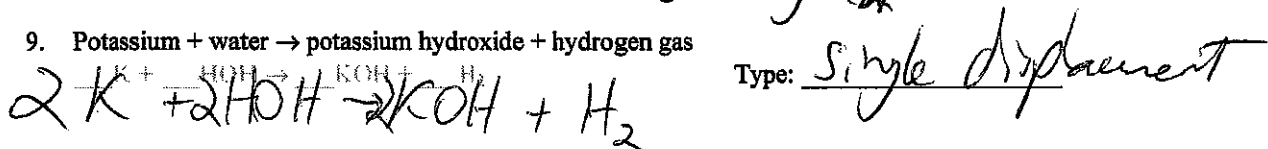
7. Magnesium + sulfuric acid → magnesium sulfate + hydrogen gas



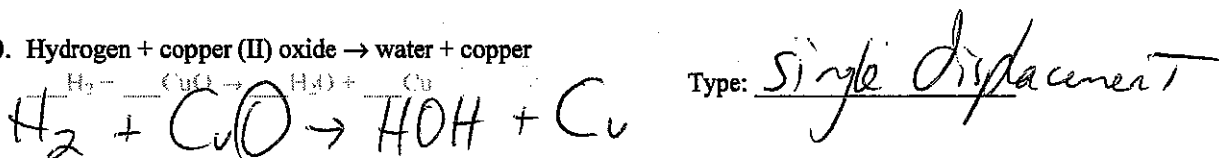
8. Silver nitrate + ammonium chloride → ammonium nitrate + silver chloride



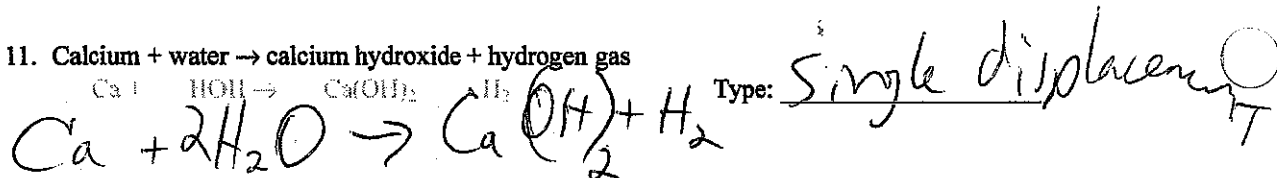
9. Potassium + water → potassium hydroxide + hydrogen gas



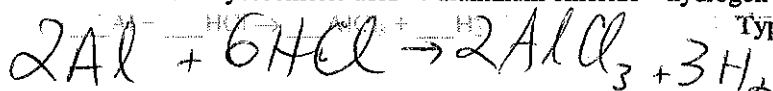
10. Hydrogen + copper (II) oxide → water + copper



11. Calcium + water → calcium hydroxide + hydrogen gas



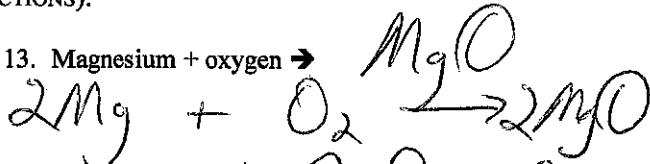
12. Aluminum + hydrochloric acid → aluminum chloride + hydrogen gas



Type: single displacement

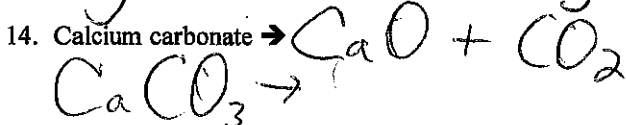
COMPLETE AND BALANCE (NOTE: "↑" MEANS MORE THAN ONE POSSIBILITY; SHOW ALL POSSIBLE REACTIONS):

13. Magnesium + oxygen →



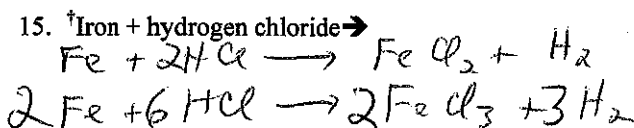
Type: Synthesis

14. Calcium carbonate →



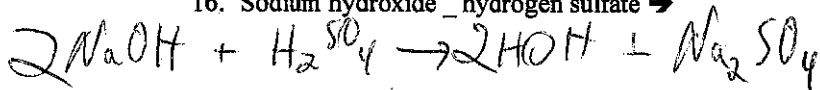
Type: decomposition

15. Iron + hydrogen chloride →



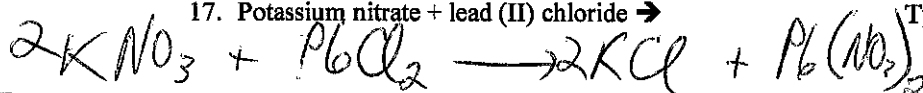
Type: single displacement

16. Sodium hydroxide + hydrogen sulfate →



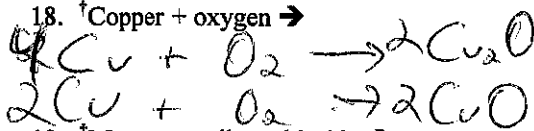
Type: neutralization
~~double displacement~~

17. Potassium nitrate + lead (II) chloride →



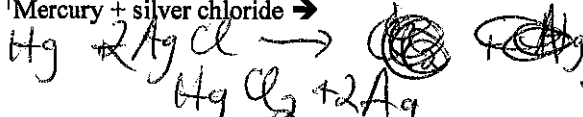
Type: double displacement

18. Copper + oxygen →



Type: Synthesis

19. Mercury + silver chloride →



Type: single displacement

20. Manganese (IV) oxide →



Type: decomposition

21. Potassium bromide + chlorine →

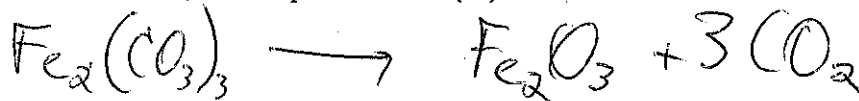


Type: single displacement

WRITE A BALANCED EQUATION FOR THE FOLLOWING REACTIONS AND IDENTIFY THE TYPE OF REACTION THAT OCCURRED (COMBUSTION, SYNTHESIS, DECOMPOSITION, SINGLE DISPLACEMENT, DOUBLE DISPLACEMENT, OR NEUTRALIZATION):

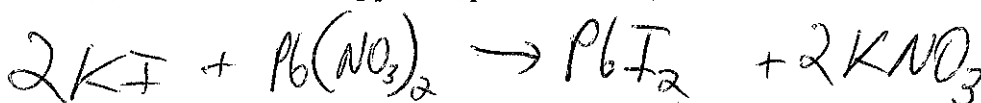
22. When iron (III) carbonate is heated, it decomposes into iron (III) oxide and carbon dioxide.

decomposition

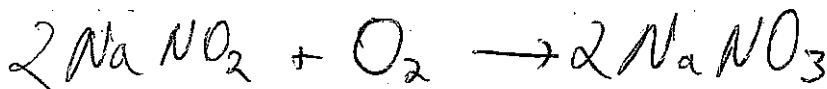


23. The bright yellow precipitate, lead (II) iodide, is produced from the reaction between potassium iodide and lead (II) nitrate. The remaining product, potassium nitrate, is left in solution.

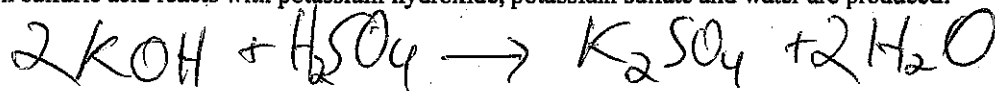
double displacement



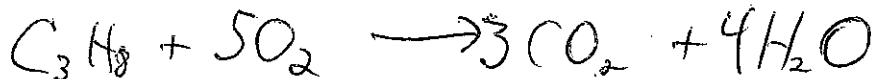
24. Sodium nitrite and oxygen gas combine to form sodium nitrate.



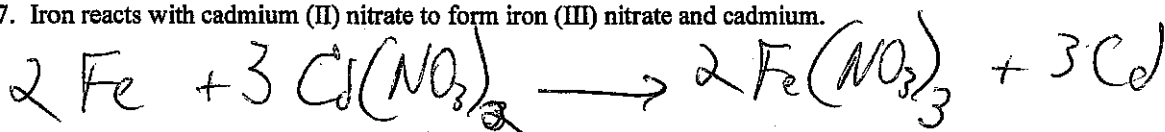
25. When sulfuric acid reacts with potassium hydroxide, potassium sulfate and water are produced.



26. ^{Propane} Pentane (C_3H_8) combusts to form carbon dioxide and water.

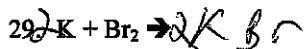
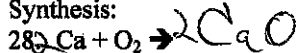


27. Iron reacts with cadmium (II) nitrate to form iron (III) nitrate and cadmium.

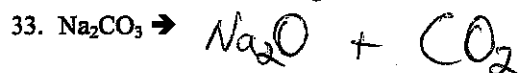
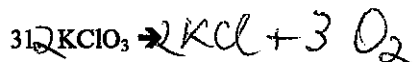
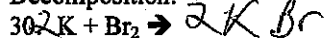


PREDICT THE PRODUCTS OF THE FOLLOWING REACTION AND THEN BALANCE THE EQUATIONS:

Synthesis:



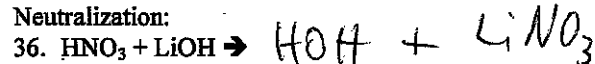
Decomposition:



Double Displacement:



Neutralization:



BALANCE THE FOLLOWING EQUATION(S) FOR EXTRA CREDIT:

