Define the following vocabulary:

- Periodic Law
- Period
- Group

Write down 2 things from the following sections:

- Recognizing a Pattern
  - Mendeleev was able to predict new elements.
  - A few elements did not fit the pattern.

- Changing the Arrangement
  - Elements become less metallic across each period.
  - Elements in a group have similar properties.
Preview of Section 5.2 “Explore the Periodic Table”

Define the following vocabulary:

- Ion
- Metal
- Nonmetal
- Semiconductor

Write down 2 things from the following sections:

- The Role of Electrons
  - Valence electrons account for similar properties.
  - An element’s location in the periodic table is related to electron arrangement.

- Ion Formation
  - Group 1 elements form positive ions.
  - Group 17 elements form negative ions.

- How Are Elements Classified?
  - Elements in each category have similar properties.
Due Date:

Preview of Section 5.3 “Families of Elements”

Name:_________________

Define the following vocabulary:

- Alkali Metal
- Alkaline–Earth Metal
- Transition Metal
- Noble Gas
- Halogen

Write down 2 things from the following sections:

- Classifying Elements Further

- Metals
  - The alkali metals are very reactive.
  - Alkaline–earth metals form compounds that are formed in limestone and in the human body.
  - Transition metals are in the middle of the periodic table.
  - Some elements are synthetic.

- Nonmetals
Due Date:

- The noble gases are relatively inert.

- The halogens combine easily with metals to form salts.

- Nonmetals and their compounds are plentiful on Earth.

- Carbon can form many compounds.

- Semiconductors

  - Hydrogen is in a class by itself.